

## MOLE CALCULATIONS 2

Titanium is extracted from titanium chloride as shown. Calculate the mass of sodium needed to react with 10 g of titanium chloride.	$TiCl_4 + 4 Na \rightarrow Ti + 4 NaCl$
Calculate the mass of oxygen needed to react 50 g of magnesium to form magnesium oxide.	2 Mg + $O_2 \rightarrow 2$ MgO
What mass of ethanol could burn in100 g of oxygen?	$C_2H_5OH + 3 O_2 \rightarrow 2 CO_2 + 3 H_2O$
What mass of hydrogen is formed when 2 g of magnesium reacts with sulfuric acid?	$Mg + H_2SO_4 \rightarrow MgSO_4 + H_2$
What was a following was to with 50 m of ablasing to form	0.41 . 0.01 . 0.4101
What mass of aluminium reacts with 50 g of chlorine to form aluminium chloride?	$2 \text{ Al} + 3 \text{ Cl}_2 \rightarrow 2 \text{ AlCl}_3$
6.15 a of hydratod magnosium gulfato docompose to form 2.00 a of	Macco villo i Macco il villo
6.15 g of hydrated magnesium sulfate decompose to form 3.00 g of anhydrous magnesium sulfate on heating. Calculate the formula mass of hydrated magnesium sulfate and the value of $x$ .	$MgSO_4.xH_2O \rightarrow MgSO_4 + x H_2O$