



# IONIC FORMULAS 3

positive ions				negative ions			
aluminium	Al <sup>3+</sup>	lead	Pb <sup>2+</sup>	bromide	Br <sup>-</sup>	iodide	I <sup>-</sup>
ammonium	NH <sub>4</sub> <sup>+</sup>	lithium	Li <sup>+</sup>	carbonate	CO <sub>3</sub> <sup>2-</sup>	nitrate	NO <sub>3</sub> <sup>-</sup>
barium	Ba <sup>2+</sup>	magnesium	Mg <sup>2+</sup>	chloride	Cl <sup>-</sup>	oxide	O <sup>2-</sup>
calcium	Ca <sup>2+</sup>	potassium	K <sup>+</sup>	fluoride	F <sup>-</sup>	sulfate	SO <sub>4</sub> <sup>2-</sup>
copper	Cu <sup>2+</sup>	silver	Ag <sup>+</sup>	hydrogencarbonate	HCO <sub>3</sub> <sup>-</sup>	sulphide	S <sup>2-</sup>
hydrogen	H <sup>+</sup>	sodium	Na <sup>+</sup>	hydroxide	OH <sup>-</sup>		
iron (II)	Fe <sup>2+</sup>	zinc	Zn <sup>2+</sup>				
iron (III)	Fe <sup>3+</sup>						

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|------------------------------|-------|----------------------------------|-------|
| 1) copper carbonate          | ..... | 11) sodium sulfate               | ..... |
| 2) lithium oxide             | ..... | 12) barium sulfide               | ..... |
| 3) silver bromide            | ..... | 13) magnesium nitrate            | ..... |
| 4) zinc nitrate              | ..... | 14) hydrogen fluoride            | ..... |
| 5) ammonium carbonate        | ..... | 15) aluminium nitrate            | ..... |
| 6) aluminium sulfate         | ..... | 16) iron (III) hydrogencarbonate | ..... |
| 7) iron (II) oxide           | ..... | 17) strontium nitrate            | ..... |
| 8) iron (III) hydroxide      | ..... | 18) chromium (IV) oxide          | ..... |
| 9) lead iodide               | ..... | 19) potassium selenide           | ..... |
| 10) sodium hydrogencarbonate | ..... | 20) rubidium oxide               | ..... |